

# Naming Chemical Compounds Worksheet

Name the following *ionic* compounds:

- 1) NaBr \_\_\_\_\_
- 2) CaO \_\_\_\_\_
- 3) Li<sub>2</sub>S \_\_\_\_\_
- 4) MgBr<sub>2</sub> \_\_\_\_\_
- 5) Pu(OH)<sub>3</sub> \_\_\_\_\_
- 6) Hg(CN)<sub>2</sub> \_\_\_\_\_
- 7) Mo(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>) \_\_\_\_\_
- 8) Cr(CrO<sub>4</sub>)<sub>3</sub> \_\_\_\_\_
- 9) WO<sub>2</sub> \_\_\_\_\_
- 10) Mg(ClO)<sub>2</sub> \_\_\_\_\_

Write the formulas for the following *ionic* compounds:

11. potassium dichromate \_\_\_\_\_
12. gold(III) oxide \_\_\_\_\_
13. aluminum periodate \_\_\_\_\_
14. barium nitrite \_\_\_\_\_
15. silver carbonate \_\_\_\_\_
16. lithium sulfite \_\_\_\_\_
17. zinc hydrogen carbonate \_\_\_\_\_
18. iron(III) hydroxide \_\_\_\_\_
19. ammonium phosphate \_\_\_\_\_
20. copper(II) bromite \_\_\_\_\_
  
- 21) SO<sub>3</sub> \_\_\_\_\_
- 22) N<sub>2</sub>S \_\_\_\_\_
- 23) BF<sub>3</sub> \_\_\_\_\_
- 24) P<sub>2</sub>Br<sub>4</sub> \_\_\_\_\_
- 25) SiO<sub>2</sub> \_\_\_\_\_
- 26) SF<sub>6</sub> \_\_\_\_\_
- 27) NO<sub>2</sub> \_\_\_\_\_
- 28) nitrogen trichloride \_\_\_\_\_
- 29) dinitrogen trioxide \_\_\_\_\_
- 30) phosphorus pentafluoride \_\_\_\_\_
- 31) diboron tetrahydride \_\_\_\_\_

# Naming Chemical Compounds - Answers

Name the following *ionic* compounds:

- 1) NaBr                    **sodium bromide**
- 2) CaO                    **calcium oxide**
- 3) Li<sub>2</sub>S                    **lithium sulfide**
- 4) MgBr<sub>2</sub>                **magnesium bromide**
- 5) Be(OH)<sub>2</sub>              **beryllium hydroxide**

Write the formulas for the following *ionic* compounds:

- 6) potassium iodide        **KI**
- 7) magnesium oxide        **MgO**
- 8) aluminum chloride      **AlCl<sub>3</sub>**
- 9) sodium nitrate          **NaNO<sub>3</sub>**
- 10) calcium carbonate      **CaCO<sub>3</sub>**
- 11) lithium sulfate         **Li<sub>2</sub>SO<sub>4</sub>**
- 12) beryllium phosphide    **Be<sub>3</sub>P<sub>2</sub>**
- 13) magnesium hydroxide   **Mg(OH)<sub>2</sub>**
- 14) sodium phosphate      **Na<sub>3</sub>PO<sub>4</sub>**
- 15) aluminum carbonate    **Al<sub>2</sub>(CO<sub>3</sub>)<sub>3</sub>**
- 16) calcium chloride       **CaCl<sub>2</sub>**
- 17) sodium cyanide         **NaCN**
- 18) aluminum oxide         **Al<sub>2</sub>O<sub>3</sub>**
- 19) magnesium acetate      **Mg(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub>**
- 20) ammonium chloride      **NH<sub>4</sub>Cl**

Write the names of the following *covalent* compounds:

- 21)  $\text{SO}_3$       **sulfur trioxide**
- 22)  $\text{N}_2\text{S}$       **dinitrogen sulfide**
- 23)  $\text{PH}_3$       **phosphorus trihydride**
- 24)  $\text{BF}_3$       **boron trifluoride**
- 25)  $\text{P}_2\text{Br}_4$       **diphosphorus tetrabromide**
- 26)  $\text{CO}$       **carbon monoxide**
- 27)  $\text{SiO}_2$       **silicon dioxide**
- 28)  $\text{SF}_6$       **sulfur hexafluoride**
- 29)  $\text{NH}_3$       **ammonia**
- 30)  $\text{NO}_2$       **nitrogen dioxide**

Write the formulas of the following *covalent* compounds:

- 31) nitrogen trichloride       **$\text{NCl}_3$**
- 32) boron carbide       **$\text{BC}$**
- 33) dinitrogen trioxide       **$\text{N}_2\text{O}_3$**
- 34) phosphorus pentafluoride       **$\text{PF}_5$**
- 35) methane       **$\text{CH}_4$**
- 36) sulfur dibromide       **$\text{SBr}_2$**
- 37) diboron tetrahydride       **$\text{B}_2\text{H}_4$**
- 38) oxygen difluoride       **$\text{OF}_2$**
- 39) carbon disulfide       **$\text{CS}_2$**
- 40) nitrogen monoxide       **$\text{NO}$**